

AD-A162 315

VIBRATIONAL DEPHASING AND ENERGY RELAXATION OF
ADMOLECULES BY PHONONS(U) STATE UNIV OF NEW YORK AT
BUFFALO DEPT OF CHEMISTRY M HUTCHINSON ET AL DEC 85

1/1

UNCLASSIFIED

UBUFFALO/DC-85-TR-2 N00014-86-K-0043

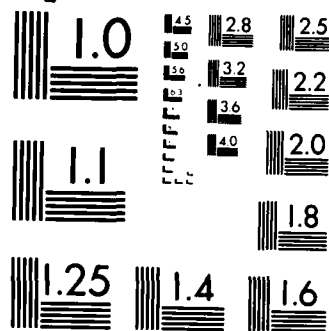
F/G 7/4

NL

END

FORMED

DTIC



MICROCOPY RESOLUTION TEST CHART
NATIONAL BUREAU OF STANDARDS 1963-A

12

Contract N00014-86-K-0043

Vibrational Dephasing and Energy Relaxation of Admolecules by Phonons

Michael Hutchinson and Thomas F. George

in

DTIC
ELECTE
DEC 16 1985

December 1985

This document has been approved for public release and sale; its distribution is unlimited.

5 12 16 114

UNCLASSIFIED

SECURITY CLASSIFICATION OF THIS PAGE

REPORT DOCUMENTATION PAGE

1a. REPORT SECURITY CLASSIFICATION Unclassified		1b. RESTRICTIVE MARKINGS	
2a. SECURITY CLASSIFICATION AUTHORITY		3. DISTRIBUTION/AVAILABILITY OF REPORT Approved for public release; distribution unlimited	
2b. DECLASSIFICATION/DOWNGRADING SCHEDULE			
4. PERFORMING ORGANIZATION REPORT NUMBER(S) UBUFFALO/DC/85/TR-2		5. MONITORING ORGANIZATION REPORT NUMBER(S)	
6a. NAME OF PERFORMING ORGANIZATION Depts. Chemistry & Physics State University of New York		6b. OFFICE SYMBOL (If applicable)	
7a. NAME OF MONITORING ORGANIZATION		7b. ADDRESS (City, State and ZIP Code) Chemistry Program 800 N. Quincy Street Arlington, Virginia 22217	
8a. ADDRESS (City, State and ZIP Code) Fronczak Hall, Amherst Campus Buffalo, New York 14260		8b. ADDRESS (City, State and ZIP Code) Chemistry Program 800 N. Quincy Street Arlington, Virginia 22217	
8a. NAME OF FUNDING/SPONSORING ORGANIZATION Office of Naval Research		8b. OFFICE SYMBOL (If applicable)	
9. PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER Contract N00014-86-K-0043		10. SOURCE OF FUNDING NOS.	
10. SOURCE OF FUNDING NOS.		11. TITLE Vibrational Dephasing and Energy Relaxation of Admolecules by Phonons	
12. PERSONAL AUTHOR(S) Michael Hutchinson and Thomas F. George		13a. TYPE OF REPORT Interim Technical	
13b. TIME COVERED FROM _____ TO _____		14. DATE OF REPORT (Yr., Mo., Day) December 1985	
15. PAGE COUNT 12		16. SUPPLEMENTARY NOTATION Physics Letters A	
17. COSATI CODES		18. SUBJECT TERMS (Continue on reverse if necessary and identify by block number)	
FIELD	GROUP	SUB. GR.	
19. ABSTRACT (Continue on reverse if necessary and identify by block number)			
An investigation is made of the vibrational dephasing of a diatomic molecule adsorbed on a surface. Explicit analytic forms for the rate of dephasing by phonons are derived. For comparison, an expression for energy relaxation is given which is appropriate for OH on SiO ₂ . It is found that the dephasing rate is considerably faster for this system than the energy relaxation rate. These conclusions are compared with the results of a recent experiment.			
20. DISTRIBUTION AVAILABILITY OF ABSTRACT UNCLASSIFIED/UNLIMITED <input checked="" type="checkbox"/> SAME AS RPT. <input checked="" type="checkbox"/> DTIC USERS <input type="checkbox"/>		21. ABSTRACT SECURITY CLASSIFICATION Unclassified	
22a. NAME OF RESPONSIBLE INDIVIDUAL Dr. David L. Nelson		22b. TELEPHONE NUMBER (Include Area Code) (202) 696-4410	
22c. OFFICE SYMBOL			

VIBRATIONAL DEPHASING AND ENERGY RELAXATION OF ADMOLECULES BY PHONONS

Michael Hutchinson*
Department of Chemistry
University of Rochester
Rochester, New York 14627

Thomas F. George
Departments of Chemistry and Physics
State University of New York at Buffalo
Buffalo, New York 14260

Abstract

An investigation is made of the vibrational dephasing of a diatomic molecule adsorbed on a surface. Explicit analytic forms for the rate of dephasing by phonons are derived. For comparison, an expression for energy relaxation is given which is appropriate for OH on SiO₂. It is found that the dephasing rate is considerably faster for this system than the energy relaxation rate. These conclusions are compared with the results of a recent experiment.

* Present address: Department of Radiology, Health Sciences Center, University of Colorado, Denver, Colorado 80262

1. Introduction

The ways in which molecular vibrations relax at surfaces are of considerable interest in dynamical studies, not least for investigating the possibility of bond-selective chemistry. Recently,¹ an experiment was reported which has for the first time directly measured the rate of vibrational energy relaxation of an adsorbed species (OH on SiO₂, $v=1 \rightarrow v=0$). Knowledge of the total width of the absorption line² then allows an estimate of the pure dephasing rate through the relation $1/T_2 = 2/T_1 + 1/T_2'$, where T_1 and T_2' are the energy and pure phase relaxation times, respectively. There are undoubtedly many relaxation pathways, including direct coupling of the molecular mode to the phonons and formation of electron-hole pairs.³ Silica, however, is a substrate in which the latter mechanism can be effectively ruled out.

We advance here a treatment of pure vibrational dephasing of an adsorbed diatomic molecule, via direct vibration-phonon coupling. For the sake of comparison, we give an expression for a T_1 pathway appropriate for OH/SiO₂. The model involves an anharmonic oscillator (A-B) coupled through an anharmonic bond (S-A) to a linear chain of harmonic oscillators. We show that phase relaxation is about 10^{-10^2} times faster for this system than energy relaxation, in broad agreement with experiment. We also find the coupling to low frequency librational modes to be unnecessary for explaining experimental results. In this regard we feel it worth emphasizing the advantages of a linear chain model when the ad molecule is aligned end-on at the surface. The assumption is that the dynamics of ad molecule and surface is dominated by the motion of the whole system in the z-direction. Such a model has recently been successfully applied to the study of the desorption of an adatom.⁴ The incorrect notion, that a linear chain model is one-dimensional in every respect, is quite widespread. In

fact, the density of states, a quantity needed for the calculation of rate information, is obtained from a fully three-dimensional (Debye) treatment, in which the phonon modes are assumed to form an isotropic continuum. The only real assumption apart from this is that the molecular dynamics is dominated by motion in a particular direction. Provided this is the case, such models are expected to give the correct order of magnitude.

2. Model

The details of the model are as follows. The momentum and position normal coordinates of the diatom-surface system are denoted by (p_1, q_1) and (p_2, q_2) . When the frequencies of the bonds are sufficiently different, the motion of the center of mass of AB relative to the surface may be identified as one normal mode, and the vibration of A relative to B as the other. Thus we may make the identities $q_1 = z$ and $q_2 = x$. We shall focus on these position coordinates for the Taylor expansion of the potential, although in fact the nonadiabatic terms due to p_1 and p_2 give contributions of a similar order of magnitude.⁵ The first term, in the Taylor expansion of the interaction potential, which gives rise to pure dephasing (i.e., the leading term in the expansion which is diagonal in the vibrational mode of the admolecule) is the quartic $k_{22}x^2$ term.⁶

In the theory of spectral lineshape, the width of a line is given by the imaginary part of the Zwanzig memory tetradic $\langle T(\omega_{fi} + i\epsilon) \rangle_{fi, fi}$,⁷ where ϵ is an infinitesimal quantity and the transition is being made between the initial (i) and final (f) state of the system. This may be expanded in powers of the interaction Liouville operator. In the Markov limit (short memory, T is constant) and in first order, there are two kinds of term: one represents contributions to the linewidth from inelastic scattering out of levels f and i , the other represents elastic transitions. The latter are identified with the "pure dephasing" interaction and an associated rate $1/T_2'$ given by:⁸

$$\frac{1}{T_2} = \frac{\pi}{\hbar^2} \sum_{\alpha, \beta} d_{\alpha\beta} (|\langle \alpha | V^{(b)} | \beta \rangle|^2 - \langle V^{(b)} \rangle^2 \delta_{\alpha\beta}) \Delta_{f1}^{(s)} \delta(\omega_{\alpha\beta}), \quad (1a)$$

where the bath-system interaction is assumed to be of the form $V^{(b)} V^{(s)}$, in obvious notation.

$$\Delta_{f1}^{(s)} = \sum_m \langle f | V^{(s)} | f \rangle_m - \sum_m \langle i | V^{(s)} | i \rangle_m \quad (2)$$

is the relative displacement of the levels i and f of the molecule due to the anharmonic interaction with the bath; d is the canonical density operator for the bath, which is assumed to be diagonal and therefore describes the bath at thermal equilibrium. The indices α and β represent entire states of the bath. Put another way, Eq. (1a) represents the thermally averaged rate of Rayleigh scattering of phonons by the system (admolecule).

On the other hand, the inelastic rates are given, in the Markov limit, by the Golden Rule expression:

$$\frac{1}{T_1} = \frac{\pi}{\hbar^2} \sum_{\alpha, \beta} \sum_f d_{\beta\beta} |\langle \alpha | V^{(b)} | \beta \rangle|^2 |V_{f1}^{(s)}|^2 \delta(\omega_{\alpha\beta} - \omega_{if}), \quad (1b)$$

which is essentially the thermally averaged rate of phonon absorption and emission by the system.

In order to introduce the bath-system coupling, we first evaluate the dynamics of AB moving with respect to a frozen surface atom S (i.e., the surface is represented by an atom of infinite mass). The influence of the thermal fluctuations of the lattice is then introduced by allowing the surface to displace by an amount u , which is fluctuating with time. Moreover, the assumption of Markovian behavior implies $u(t)$ is uncorrelated with $u(t')$, $t \neq t'$. We replace z with $z-u$ and express u in terms of a Fourier sum over the normal coordinates of the lattice:⁹

$$u(t) = \left(\frac{\hbar}{2NM}\right)^{1/2} \sum_k \omega_k^{-1/2} (a_k^\dagger e^{-i\omega_k t} + a_k e^{i\omega_k t}), \quad (3)$$

where there are N atoms in the chain, each of mass M , and a_k^\dagger, a_k are the creation and annihilation operators for mode k , which has frequency ω_k .

First, we consider the dephasing of a two-level molecule attached to a surface. By making a Taylor expansion of the interaction it is possible to pick out terms which contribute to the process of interest. The first term in the interaction which gives rise to pure dephasing is $k_{22} u^2 x^2$, and therefore

$$V^{(b)}_{V(s)} = k_{202} u^2 x^2. \quad (4)$$

Here, the subscripts represent the degree of excitation of phonons, adbond and molecule, respectively. Substituting Eqs. (3) and (4) into Eq. (1a) yields

$$\frac{1}{T_2'} = \frac{\pi k_{22}^2}{4M^2 N^2} \sum_{k,k'} d_{\beta\beta} (|\langle \alpha | a_k^\dagger a_{k'} + a_k a_{k'}^\dagger | \beta \rangle|^2 - \bar{n}_k^2 \delta_{kk'}) \Delta_{10}^{(s)} \delta(\omega_k - \omega_{k'}). \quad (5a)$$

Here, \bar{n}_k is the thermal average of n_k , and is equal to $(e^\theta - 1)^{-1}$, with $\theta = \hbar\omega_k/k_B T$.

Converting the sum to an integral gives

$$\frac{1}{T_2'} = \frac{\pi k_{22}^2}{4M^2 N^2} \Delta_{10}^{(s)} \int d\omega_k d\omega_{k'} \frac{\rho(\omega_k) \rho(\omega_{k'})}{\omega_k \omega_{k'}} \bar{n}_k (\bar{n}_{k'} + 1), \quad (5b)$$

where $\rho(x)$ is the density of modes of frequency x . We now invoke the Debye approximation, in which the phonon states of a solid form an isotropic continuum with a density of states given by $\rho(\omega_k) = \frac{3\omega_k^2}{\omega_D^3} N$, where ω_D is the Debye frequency of the solid. This yields:

$$\frac{1}{T_2'} = \frac{9\pi k_{22}^2}{5M^2 \omega_D^6} \Delta_{10}^{(s)} I, \quad (6a)$$

where

$$I = \int_0^{\omega_D} d\omega_k (\bar{n}_k^2 + \bar{n}_k) \omega_k^2. \quad (6b)$$

In order to evaluate $\Delta_{10}^{(s)}$ in Eq. (6a), we take the states $|1\rangle_m$ and $|0\rangle_m$ to be states of a harmonic oscillator, when

$${}_m \langle j | x^2 | j \rangle_m = \frac{2j+1}{2\mu_{AB}\omega_{AB}} \hbar \quad (7)$$

Here, μ_{AB} is the reduced mass of the admolecule and ω_{AB} its fundamental frequency. Moreover, we can define two limiting cases for which analytic solutions to Eq. (6b) are possible. In the limit of low temperature, $\theta \gg 1$, and ω_D may be taken to be infinity. In the limit of high temperature, $\theta \ll 1$ and the exponentials in the expression for \bar{n}_k may be expanded, while retaining only the first term. Then we arrive at the following equations for the pure dephasing rate:

$$\frac{1}{T_2^1} = \frac{3\pi^3 k_{02}^2 k_B^3}{4M^2 \mu_{AB}^2 \omega_{AB}^2 \omega_D^6} T^3 \quad k_B T \ll \hbar \omega_D \quad (8a)$$

$$\frac{1}{T_2^1} = \frac{9\pi k_{02}^2 k_B^2}{4M^2 \mu_{AB}^2 \omega_{AB}^2 \omega_D^5} T^2 \quad k_B T \gg \hbar \omega_D \quad (8b)$$

Consider now the energy relaxation of A-B. In the case of OH/SiO₂, $\omega_D = 340$ cm⁻¹, $\omega_{AS} = 970$ cm⁻¹ and $\omega_{AB} = 3700$ cm⁻¹. Therefore, the first important interaction giving rise to energy relaxation is the sextic term, $(z-u)^5 x$, and in particular $uz^4 x$ which destroys one phonon, creates four quanta of adbond excitation and destroys the quantum of vibrational excitation in OH. Moreover we note that a direct one phonon deactivation of AB is not allowed by energy conservation, and that it would require at least eleven phonons.

We now identify eigenstates of the system as products of adbond states and molecular states. Using Eq. (1b) together with the Debye density of states yields

$$\frac{1}{T_1} = \frac{3\pi\gamma^2}{2M\hbar\omega_D^3} \bar{n}(\omega_p) \omega_p \Theta(\omega_p - \omega_D) \quad (9a)$$

where

$$\gamma = k_{141} \langle 4|z^4|0 \rangle_a \langle 0|x|1 \rangle_m, \quad (9b)$$

and $\omega_p = 4\omega_{SA} - \omega_{AB}$; k_{141} is the coefficient of uz^4x . For simplicity we have taken the eigenstates of the adbond to be harmonic oscillator states. This is clearly a cruder approximation than making AB harmonic, especially given the level of excitation of SA (four quanta). However, there is little detailed information about the silica-hydroxyl bond, and we are in any case interested only in making an estimate of the rate of energy relaxation for this system. The introduction of anharmonicities into the SA bond will make the calculation only slightly more complicated. (Clearly, the interaction Hamiltonian depends only on the anharmonicity of the bond!)

It remains to determine k_{202} and k_{141} . Obviously, we must here take a proper account of the anharmonicity of the adbond, SA; indeed we shall make the assumption that this is the principal contributor to the coupling between the adbond and the molecular band AB. The validity of this assumption will be assessed at a later time.¹⁰ Further we make the assumption that z , the distance between the surface and the center-of-mass of AB, is given approximately by the distance SA. This assumption is expected to be reasonable when A is much more massive than B (as in OH). Therefore, the motion of the AB bond gives rise to a modified Morse potential for SA in which the distance parameter is $z - \sigma x$, where σ is the ratio (M_B/M_A). A simple expansion of the Morse potential around the equilibrium distance then gives terms containing various combinations of powers of z and x . We are interested in the coefficient of the terms z^2x^2 and z^5x . Next we replace z with $z - u$ and pick out the coefficients of u^2x^2 and uz^4x . We obtain finally

$$k_{202} = \frac{35}{6} a^4 \sigma^2 D \quad (10a)$$

$$k_{141} = \frac{31}{18} a^6 \sigma D, \quad (10b)$$

where a is the strength parameter of the AS Morse potential and D is the well depth. The parameter a may be written as $(\mu_{AS}/3D)^{1/2}$. To the best of our knowledge, D is an unknown quantity for OH/SiO₂, but we shall take it to be 1 eV, noting that both energy and phase relaxation depend upon D^{-2} . Thus changing D by a factor of 2 will result in a change in rate of a factor of 4. Finally, we use a Debye temperature for SiO₂ of 485 K and take the molecular weight, M , of the SiO₂ unit cell to be 60. From Eqs. (8a) and (9a) we now obtain $T_2' = 4\text{ps}$ and $T_1 = 2 \times 10^4 \text{ ps}$, corresponding to widths of 1.3 cm^{-1} and $2.6 \times 10^{-4} \text{ cm}^{-1}$, respectively. It is immediately clear that the calculated value for T_1 is about two orders of magnitude larger than the experimental value.¹ However, the value of T_2' is in much better agreement with the available information on this system,¹¹ where it is believed that the pure dephasing width $\leq 1 \text{ cm}^{-1}$.

The large discrepancy between the observed and calculated values of T_1 is very interesting, particularly since one of the approximations involved the assumption of a harmonic adbond for the calculation of k_{141} in Eq. (9b). Removal of this assumption will undoubtedly increase the estimate of T_1 still further. It is unlikely that the approximation either of the Debye model or of the linear chain dynamics are very important here. Of much greater importance is the crudity of our use of a Morse potential to estimate anharmonicities, particularly as (for T_1) the relevant terms are sextic.

However, the discrepancy is sufficiently large that we may tentatively postulate a different mechanism for the energy relaxation. The reason the mechanism given here leads to such long relaxation times is that the vibrational

quantum of the molecular bond is so large that energy transfer requires several quanta of excitation of other modes, and this is a process requiring at least a sextic anharmonicity. The direct excitation of the substrate requires at least eleven phonons and may be presumed to be insignificant. It is possible that the OH couples to another kind of high-frequency mode, such as the Si-OH bend, although this mode is likely to be of a somewhat lower frequency than the Si-OH stretch we have considered. An interesting possibility is the direct energy exchange between adjacent hydroxyl groups on the surface. Even for weak coupling, such a pathway may be favored because the vibrational energy mismatch is zero. However, the irreversibility of energy relaxation demands several adjacent, anharmonically-coupled hydroxyl groups. In this regard, therefore, it will be interesting to see if T_1 increases with decreasing coverage.

As far as pure dephasing is concerned, the influence of an adjacent hydroxyl may be crudely modelled in terms of a dipole-dipole interaction.¹³ Since the dipole moment depends upon vibrational state, such an interaction gives rise to a relative displacement of vibrational levels, which leads to a dephasing. For H_2O the vibrational dependence of the dipole is known and the dephasing contribution at room temperature may be estimated to be about $.01 \text{ cm}^{-1}$. Assuming things are similar for OH, we can therefore rule out this contribution to the pure dephasing, since the calculated phonon-mediated contribution is two orders of magnitude faster. The calculation of the contribution of this admolecule-admolecule interaction to energy relaxation is less straightforward, and will be possible only when the anharmonic interaction between adjacent hydroxyl groups is known. It is also possible that the k_{141} term is much larger than our estimate, so that energy exchange with the lattice is important. Definitive answers await further experimental information.

During preparation of this manuscript we become aware of a recent paper by Persson,¹² which is similar in spirit to the present work. There a comparison

was made between energy relaxation and dephasing of the adbond, specifically for CO on nickel. The present work, on the other hand, describes relaxation of the molecular bond. Moreover, we have for the first time presented explicit analytical forms for T_2^1 .

Conclusion

We have derived expressions for pure vibrational dephasing and for vibrational energy relaxation of a diatomic molecule on a surface. The model involves the coupling of the molecular mode to bulk phonons. This is given in terms of the anharmonicities of the molecular and ad molecular vibrations. By making an estimate of these quantities for OH on SiO_2 , we have shown that the prediction of the pure vibrational phase relaxation is in agreement with available experimental information, but the energy relaxation rate is two orders of magnitude smaller than experiment. This indicates the possibility of another mechanism in this case, and the obvious candidate is energy transfer between adjacent hydroxyl groups. A definitive answer will not be possible until the anharmonicities are known accurately. The influence of surface phonons will be assessed at a later time.¹⁰ The inclusion of such modes is expected to be important for molecules lying across the surface as opposed to end-on, as in the present case. This is because of the expected efficient coupling between the lateral motion of the surface atoms and the motion of the molecular bond.

Acknowledgments

This research was supported in part by the Office of Naval Research, the Air Force Office of Scientific Research (AFSC), United States Air Force, under Contract F49620-86-C-009, and the National Science Foundation under Grant CHE-8519053. The United States Government is authorized to reproduce and distribute reprints for governmental purposes notwithstanding any copyright notation hereon. TFG acknowledges the Camille and Henry Dreyfus Foundation for a Teacher-Scholar (1975-86).

References

1. E. J. Heilwell, M. P. Casassa, R. R. Cavanagh and J. C. Stephenson, *J. Chem. Phys.* 81, 2856 (1984).
2. W. Hertl and M. L. Hair, *J. Phys. Chem.* 72, 4676 (1968).
3. B. N. J. Persson and R. Ryberg, *Phys. Rev. Lett.* 48, 549 (1982); H. Metiu and W. E. Palke, *J. Chem. Phys.* 69, 2574 (1978).
4. S. Efrima, C. Jedrzejek, K. F. Freed, E. Hood and H. Metiu, *J. Chem. Phys.* 79, 2436 (1983); E. Hood, C. Jedrzejek, K. F. Freed and H. Metiu, *J. Chem. Phys.* 81, 3277 (1984).
5. J. K. G. Watson, *Mol. Phys.* 19, 465 (1970); T. Oka, *J. Chem. Phys.* 57, 5410 (1967).
6. This has also been noted by J. W. Gadzuk and A. C. Luntz, *Surf. Sci.* 144, 429 (1984).
7. R. Zwanzig, *J. Chem. Phys.* 33, 1338 (1960).
8. U. Fano, in Lectures on the Many-Body Problem, vol. 2, ed. by E. R. Cianello (Academic, New York, 1964). A concise account of the relaxation theory of Fano is given by M. J. Burns, W. K. Liu and A. H. Zewail, in Spectroscopy and Excitation Dynamics of Condensed Molecular Systems, ed. by V. M. Agranovich and R. M. Hochstrasser (North-Holland, Amsterdam, 1983).
9. A. A. Maradudin, E. W. Montroll, G. H. Weiss and I. P. Ipatova, Theory of Lattice Dynamics in the Harmonic Approximation (Academic, New York, 1971), Chapt. 1.
10. M. Hutchinson and T. F. George, to be published.
11. T_1 is determined as 150 ps in reference 1. The experimental linewidth of $4-8 \text{ cm}^{-1}$ (corresponding to 1.3-0.65 ps) for OH at the silica/vacuum interface is known to have an inhomogeneous component. Therefore, this must represent an upper bound on the dephasing contribution (J. C. Stephenson, private communication). It is believed that the width due to dephasing is $\leq 1 \text{ cm}^{-1}$, which corresponds to $T_2' \geq 5 \text{ ps}$. Regarding our use of the low temperature formula when the temperature is actually a significant fraction of the Debye temperature (61%): we estimate that this will result in a slight (~20%) positive error in the calculation of the pure dephasing rate. Hence our estimate of $\sim 1 \text{ cm}^{-1}$ should be regarded as an upper bound.
12. B. N. J. Persson, *J. Phys. C* 17, 4741 (1984).
13. M. Hutchinson, X. Y. Huang, K. C. Liu, K. T. Lee and T. F. George, in Laser Chemical Processing of Semiconductor Devices, Extended Abstracts of the Proceedings of Symposium B of the 1984 Fall Meeting of the Materials Research Society (Boston, Massachusetts, 1984), pp. 94-95.

DL/413/83/01
GEN/413-2

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	<u>No. Copies</u>		<u>No. Copies</u>
Office of Naval Research Attn: Code 413 800 N. Quincy Street Arlington, Virginia 22217	2	Dr. David Young Code 334 NORDA NSTL, Mississippi 39529	1
Dr. Bernard Douda Naval Weapons Support Center Code 5042 Crane, Indiana 47522	1	Naval Weapons Center Attn: Dr. Ron Atkins Chemistry Division China Lake, California 93555	1
Commander, Naval Air Systems Command Attn: Code 310C (H. Rosenwasser) Washington, D.C. 20360	1	Scientific Advisor Commandant of the Marine Corps Code RD-1 Washington, D.C. 20380	1
Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko Port Hueneme, California 93401	1	U.S. Army Research Office Attn: CRD-AA-IP P.O. Box 12211 Research Triangle Park, NC 27709	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	12	Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 19112	1
DTNSRDC Attn: Dr. G. Bosmajian Applied Chemistry Division Annapolis, Maryland 21401	1	Naval Ocean Systems Center Attn: Dr. S. Yamamoto Marine Sciences Division San Diego, California 91232	1
Dr. William Tolles Superintendent Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375	1	Dr. David L. Nelson Chemistry Division Office of Naval Research 800 North Quincy Street Arlington, Virginia 22217	1

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. G. A. Somorjai
Department of Chemistry
University of California
Berkeley, California 94720

Dr. J. Murday
Naval Research Laboratory
Surface Chemistry Division (6170)
455 Overlook Avenue, S.W.
Washington, D.C. 20375

Dr. J. B. Hudson
Materials Division
Rensselaer Polytechnic Institute
Troy, New York 12181

Dr. Theodore E. Madey
Surface Chemistry Section
Department of Commerce
National Bureau of Standards
Washington, D.C. 20234

Dr. J. E. Demuth
IBM Corporation
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Dr. M. G. Lagally
Department of Metallurgical
and Mining Engineering
University of Wisconsin
Madison, Wisconsin 53706

Dr. R. P. Van Duyne
Chemistry Department
Northwestern University
Evanston, Illinois 60637

Dr. J. M. White
Department of Chemistry
University of Texas
Austin, Texas 78712

Dr. D. E. Harrison
Department of Physics
Naval Postgraduate School
Monterey, California 93940

Dr. W. Kohn
Department of Physics
University of California, San Diego
La Jolla, California 92037

Dr. R. L. Park
Director, Center of Materials
Research
University of Maryland
College Park, Maryland 20742

Dr. W. T. Peria
Electrical Engineering Department
University of Minnesota
Minneapolis, Minnesota 55455

Dr. Keith H. Johnson
Department of Metallurgy and
Materials Science
Massachusetts Institute of Technology
Cambridge, Massachusetts 02139

Dr. S. Sibener
Department of Chemistry
James Franck Institute
5640 Ellis Avenue
Chicago, Illinois 60637

Dr. Arnold Green
Quantum Surface Dynamics Branch
Code 3817
Naval Weapons Center
China Lake, California 93555

Dr. A. Wold
Department of Chemistry
Brown University
Providence, Rhode Island 02912

Dr. S. L. Bernasek
Department of Chemistry
Princeton University
Princeton, New Jersey 08544

Dr. P. Lund
Department of Chemistry
Howard University
Washington, D.C. 20059

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. F. Carter
Code 6132
Naval Research Laboratory
Washington, D.C. 20375

Dr. Richard Colton
Code 6112
Naval Research Laboratory
Washington, D.C. 20375

Dr. Dan Pierce
National Bureau of Standards
Optical Physics Division
Washington, D.C. 20234

Dr. R. Stanley Williams
Department of Chemistry
University of California
Los Angeles, California 90024

Dr. R. P. Messmer
Materials Characterization Lab.
General Electric Company
Schenectady, New York 22217

Dr. Robert Gomer
Department of Chemistry
James Franck Institute
5640 Ellis Avenue
Chicago, Illinois 60637

Dr. Ronald Lee
R301
Naval Surface Weapons Center
White Oak
Silver Spring, Maryland 20910

Dr. Paul Schoen
Code 5570
Naval Research Laboratory
Washington, D.C. 20375

Dr. John T. Yates
Department of Chemistry
University of Pittsburgh
Pittsburgh, Pennsylvania 15260

Dr. Richard Greene
Code 5230
Naval Research Laboratory
Washington, D.C. 20375

Dr. L. Kesmodel
Department of Physics
Indiana University
Bloomington, Indiana 47403

Dr. K. C. Janda
California Institute of Technology
Division of Chemistry and Chemical
Engineering
Pasadena, California 91125

Dr. E. A. Irene
Department of Chemistry
University of North Carolina
Chapel Hill, North Carolina 27514

Dr. Adam Heller
Bell Laboratories
Murray Hill, New Jersey 07974

Dr. Martin Fleischmann
Department of Chemistry
Southampton University
Southampton SO9 5NH
Hampshire, England

Dr. John W. Wilkins
Cornell University
Laboratory of Atomic and
Solid State Physics
Ithaca, New York 14853

Dr. Richard Smardzewski
Code 6130
Naval Research Laboratory
Washington, D.C. 20375

Dr. H. Tachikawa
Chemistry Department
Jackson State University
Jackson, Mississippi 39217

DL/413/83/01
056/413-2

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. R. G. Wallis
Department of Physics
University of California
Irvine, California 92664

Dr. D. Ramaker
Chemistry Department
George Washington University
Washington, D.C. 20052

Dr. J. C. Hemminger
Chemistry Department
University of California
Irvine, California 92717

~~Dr. T. F. George
Chemistry Department
University of Rochester
Rochester, New York 14627~~

Dr. G. Rubloff
IBM
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Dr. Horia Metiu
Chemistry Department
University of California
Santa Barbara, California 93106

Dr. W. Goddard
Division of Chemistry
California Institute of Technology
Pasadena, California 91125

Dr. J. T. Keiser
Department of Chemistry
University of Richmond
Richmond, Virginia 23173

Dr. R. W. Plummer
Department of Physics
University of Pennsylvania
Philadelphia, Pennsylvania 19104

Dr. E. Yeager
Department of Chemistry
Case Western Reserve University
Cleveland, Ohio 41106

Dr. N. Winograd
Department of Chemistry
Pennsylvania State University
University Park, Pennsylvania 16802

Dr. Roald Hoffmann
Department of Chemistry
Cornell University
Ithaca, New York 14853

Dr. A. Steckl
Department of Electrical and
Systems Engineering
Rensselaer Polytechnic Institute
Troy, New York 12181

Dr. G. H. Morrison
Department of Chemistry
Cornell University
Ithaca, New York 14853

Dr. P. Hansma
Physics Department
University of California
Santa Barbara, California 93106

Dr. J. Baldeschwieler
California Institute of Technology
Division of Chemistry
Pasadena, California 91125

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. J. E. Jensen
Hughes Research Laboratory
3011 Malibu Canyon Road
Malibu, California 90265

Dr. J. H. Weaver
Department of Chemical Engineering
and Materials Science
University of Minnesota
Minneapolis, Minnesota 55455

Dr. W. Goddard
Division of Chemistry
California Institute of Technology
Pasadena, California 91125

Dr. A. Reisman
Microelectronics Center of North Carolina
Research Triangle Park, North Carolina
27709

Dr. M. Grunze
Laboratory for Surface Science and
Technology
University of Maine
Orono, Maine 04469

Dr. J. Butler
Naval Research Laboratory
Code 6115
Washington D.C. 20375

Dr. L. Interante
Chemistry Department
Rensselaer Polytechnic Institute
Troy, New York 12181

Dr. Irvin Heard
Chemistry and Physics Department
Lincoln University
Lincoln University, Pennsylvania 19352

Dr. K.J. Klaubunde
Department of Chemistry
Kansas State University
Manhattan, Kansas 66506

Dr. W. Knauer
Hughes Research Laboratory
3011 Malibu Canyon Road
Malibu, California 90265

Dr. C. B. Harris
Department of Chemistry
University of California
Berkeley, California 94720

Dr. F. Kutzler
Department of Chemistry
Box 5055
Tennessee Technological University
Cookeville, Tennessee 38501

Dr. D. DiLella
Chemistry Department
George Washington University
Washington D.C. 20052

Dr. R. Reeves
Chemistry Department
Rensselaer Polytechnic Institute
Troy, New York 12181

END

FILMED

1-86

DTIC